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Section Experiment 30 Advance Study
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1. write the balanced net ionic equation for the reaction between MnO_4^- ion and Fe^{2+} ion in acidic solution.
2. How many moles of Fe^{2+} ion can be

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oxidized by $1A \times 10$ moles MnO_4^- ion in
the reaction in Question 1? 3.

Solved: Section Experiment 30

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$$7 \text{ MnO}_4^- + 5 \text{ Fe}^{2+} + 8 \text{ H}^+ \rightarrow 7 \text{ Mn}^{2+} + 5 \text{ Fe} + 4 \text{ H}_2\text{O}$$

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Experiment 21 Advance Study

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Clock Reaction? 1. A student studied

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the clock reaction described in this experiment. she set up Reaction Mixture 3 by mixing 10 mL 0.010 M KI, 10 mL 0.0010 Na₂S₂O₃, 20 mL 0.040 M KBrO₃, and 10 mL M HCl. Advance Study Assignment: Rate of Reactions II: A ...

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Iodination of Acetone 1. In a reaction

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Experiment 36 Answer
involving the iodination of acetone, the following volumes were used to make up the reaction mixture: 5 mL. 4.0 M acetone+10 mL. 1.0 M HCl+10 mL.0.0050 M I₂+25 mL ...

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