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Figure 7.3.6 Absorption and Emission Spectra. Absorption of light by a hydrogen atom. (a) When a hydrogen atom absorbs a photon of light, an electron is excited to an orbit that has a higher energy and larger value of

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Lab Answers (b) Images of the emission and absorption spectra of hydrogen are shown here.

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The general formula for the hydrogen emission spectrum is

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Lab Answers given by: Where,  $n_1 = 1, 2, 3, 4 \dots$   
 $n_2 = n_1 + 1$ .  $\nu$  = wave number of electromagnetic radiation. The value  $109,677 \text{ cm}^{-1}$  is known as Rydberg constant for hydrogen. To learn more about hydrogen emission spectrum download BYJU'S - The Learning App. Read

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more: Emission spectrum and atomic spectra; Spectroscopy

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series. The various series of the spectrum of atomic hydrogen are listed in . Table 14.1. According to the Bohr theory, the Rydberg constant,  $R$ , is given to first approximation by,  $2 \ 1 \ 3 \ 2 \ 4 \ 1$ .  
097 10 4 x. nm c k me R. S=

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Equation 14.2, where  $e$  and  $m$  are the electron charge and mass, respectively,  $c$  is the velocity of light,  $h$  is Planck's constant divided by  $2\pi$ , and  $n$

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An atomic emission spectrum of hydrogen shows the following three wavelengths: 1875 nm, 1282 nm, and 1093 nm. Assign these wavelengths to transitions in the hydrogen atom. For nm = 1282 .

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Hydrogen molecules are first broken up into hydrogen atoms (hence the atomic hydrogen emission spectrum) and electrons are then promoted into higher

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The emission spectrum of hydrogen Lyman series: It is made of all the de-excitations that end up on the  $n_f = 1$  level. Infinite number of them:  $n_i = 2, 3, 4, \dots$   $n_f = 1$ . Unfortunately the Lyman series is not visible with

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in the ultraviolet region-13.6 eV  
0.0 eV E ... PHYS

1493/1494/2699: Exp. 7 -  
Spectrum of the Hydrogen Atom

Experiment 7: Spectrum of the  
Hydrogen Atom



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The emission spectrum of atomic hydrogen has been divided into a number of spectral series, with wavelengths given by the Rydberg formula. These observed spectral lines are due to the electron making transitions between two energy levels in an

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atom. The classification of the series by the Rydberg formula was important in the development of quantum mechanics. The spectral series are important in astronomical spectroscopy for detecting the presence of hydrogen and

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calculating red shifts.

[Hydrogen spectral series - Wikipedia](#)

of the spectrum of atomic hydrogen was among the strongest evidence for the validity of the "new" theory of quantum

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An atomic emission spectrum of hydrogen shows three wavelengths: 1,875 nm, 1,282 nm, and 1,093 nm. Assign these wavelengths to transitions in the hydrogen atom.

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The atomic spectra are defined as the spectrum of frequencies of electromagnetic radiation emitted or absorbed during transitions of electrons between energy levels within an atom. ... The formula was primarily presented as a generalization of the Balmer

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series for all atomic transitions of hydrogen. It is a unit of energy defined in terms of ...

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The emission spectrum of hydrogenSome of the most



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common and readily observable series have been named as shown in this image, where  $n_1$  is the ground state and  $n_2$  are excited states. The various series are named for the atomic energy level they end on ( $n_1$ ). The series limit where  $n_2$  is infinite and  $n_1$

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$n=1$  corresponds to the ionization energy of hydrogen.

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In this video we discuss emission spectrum and absorption

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spectrum of atoms & look at at various spectral series for Hydrogen atom. We know from Blackbody rad...

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When this light is passed through

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a prism the result is not a continuous spectrum, but a line spectrum. In other words, only certain wavelengths ( ) of light are observed. This observation strongly suggests that the energy absorbed/emitted by an "Lab 1: Atomic Spectrum of Hydrogen" by

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A line in the spectrum of atomic mercury has a wavelength of 258 nm. When mercury emits a

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photon of light at this wavelength, the frequency of this light is: a.  $8.61 \times 10^{-16} \text{ s}^{-1}$  b.  $7.70 \times 10^{-19} \text{ s}^{-1}$  c.  $1.16 \times 10^{15} \text{ s}^{-1}$  d.  $77.3 \text{ s}^{-1}$  e. none of these

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Atomic Hydrogen 5. The energy of an electron in the hydrogen atom is given by the equation  $E_n = -2.18 \times 10^{-18} \text{ J } (1/n^2)$  where 'n' refers to the principal energy level or shell occupied by the electron ( $n = 1, 2, 3, 4, \dots$  etc.)

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