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Equation 7.3.1 and Equation 7.3.2 are balanced chemical equations. What is different on each

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side of the equation is how the atoms are arranged to make molecules or ions. A chemical reaction represents a change in the distribution of atoms but not in the number of atoms. In this reaction, and in most chemical reactions, bonds are broken in the reactants (here, Cr–O and N–H bonds), and new bonds are formed to create the products (here, O–H and N–N bonds).

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mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples

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CHAPTER 7 REVIEW Chemical Formulas and Chemical Compounds SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Label each of the following statements as True or False: True a. If the formula mass of one molecule is x amu, the molar mass is x g/mol. False b. Samples of equal numbers of moles of two different chemicals

7 Chemical Formulas and Chemical Compounds

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Chapter 7 3 Answers Chemical As shown in Figure 7.3.1, applying a small amount of heat to a pile of orange ammonium dichromate powder results in a vigorous reaction known as the Page 4/30

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ethanol (CH₃CH₂OH) Answer: Ba, +2; F, ?1; C, 0; H, +1; O, ?2; K, +1; Cr, +6; O, ?2; Cs, +1; O, ?½; C, ?3; H, +1; C, ?1; H, +1; O, ?2; H, +1

Chapter 7.5: Types of Chemical Reactions - Chemistry ...

Answer. (a) baking soda (sodium bicarbonate) plus vinegar (acetic acid) yields carbon dioxide plus water plus sodium ion plus acetate ion. Vinegar (Acetic acid) + Baking soda (Sodium hydrogencarbonate) -> Carbon dioxide + other substances. $\text{NaHCO}_3 + \text{CH}_3\text{COOH} \rightarrow \text{NaC}_2\text{H}_3\text{O}_2 + \text{H}_2\text{O} + \text{CO}_2$

Physical and chemical changes class 7 Questions with Answers

Chapter 7: Chemical Nomenclature 7.1 Molecular Formula Practice Questions ... 7.3 Cations Practice ... 3) 2 d. NaCN Answers 1. A compound with three or more elements. 2. Same rule as naming binary ionic compounds. 3. a. Hg(NO

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Chapter 7 Chemical Reactions Answer Key

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Iron and sulfur react to form Iron (III) sulfide 2. Phosphorus and oxygen combine to form diphosphorus pentoxide 3. Water decomposes to form hydrogen and oxygen 4. Sodium reacts with water to form sodium hydroxide and hydrogen 5. Iron metal reacts with copper (II) nitrate to form iron (II) nitrate and copper metal 6.

Class Worksheet-2 Chapter 7 Name More Practice Pre ...

Balancing Chemical Equations Chapter 7 Balancing Chemical Equations Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2$

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