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2, Levenspiel 1. If the reaction is not elementary, one cannot know the order. If the reaction is

elementary, the order is 2<sup>nd</sup> in either direction.  $-r_A = k_1[A][B] - k_2[A][R]$

$R = \frac{1}{2} \frac{d[A]}{dt} = -\frac{1}{2} \frac{d[R]}{dt}$

3. The same.  $(r_A = -\frac{1}{2} \frac{d[A]}{dt})$

4. regardless of how the stoichiometric equation is written.)

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