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### Chapter 16 The Periodic Table - O Level Chemistry tuition ...

The error in print is on Page 271 under the Chapter of " Electrolysis " in Chemistry Matters for Pure Chemistry syllabus. In the Key Ideas which acts like sub-section summary, it states that " OH- ions will be discharged as oxygen gas. However, if a solution is concentrated in halide ions, the halogen gas will be produced."

### Error in Marshall Cavendish GCE O-Level Chemistry Matters ...

Chemistry Matters for GCE 'O' Level (2 ndEdition): Textbook Answers Chapter 9 © 2013 Marshall Cavendish International (Singapore) Private Limited 9.2 (c) Number of moles of C 12 gas = mass of C 12 molar mass of C 12 = 1.42 71 = 0.02 mol (d) Number of moles of sodium chloride = number of ions of sodium chloride  $6 \times 10^{23} = 6 \times 10^{22} \times 6 \times 10^{23} = 0.1 \text{ mol } 2.$

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Test Yourself 4.1 (page 74) 1. (Any four differences) Mass Weight • An amount of matter • A gravitational force • A scalar quantity (i.e. has only magnitude) • A vector quantity (i.e. has both magnitude and direction) • SI unit: kilogram (kg) • SI

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Chemistry Matters for GCE 'O' Level (2 ndEdition): Textbook Answers Chapter 10 © 2013 Marshall Cavendish International (Singapore) Private Limited 10.2 Test Yourself 10.2 (page 167) 1.  $2\text{C}_2\text{H}_6(\text{g}) + 7\text{O}_2(\text{g}) \rightarrow 4\text{CO}_2(\text{g}) + 6\text{H}_2\text{O}(\text{g})$  Volume of carbon dioxide formed = 2 ×

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